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Modeling Of Potassium Uptake in Arid Region under Water Scarcity

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Abstract— ***The objective of this study is to determine the kinetics of potassium (K^+) release from sandy soil, represents new reclaimed areas in Egypt as affected by compost and chemical fertilizers applied individually or mixture of both types in the studied soil under two irrigation regimes, i.e. 80 and 60% of calculated irrigation water requirements (IR). Four mathematical models i.e. power function or modified Freundlich equation (MFE), Elovich, parabolic diffusion and first-order models were used to describe K desorption reactions. Data indicated that according to higher correlation coefficient of determination (R2) and the lowest standard error (SE), Elovich and MFE models overall displayed the best fitted ones. The first-order rate and for less extent, parabolic diffusion equation were also described the data. The rate constants for used models indicated that all types of fertilizers applied to soil gave high and significant increase in rate of K release compared to the control (untreated soil). According to the same constants values, purely organic fertilizer applied gave the lowest rate value compared to chemical one, the rate constants values of the mixture treatments, however, gave intermediate values between chemical and organic fertilizers. Concerning the water regime applied, data showed that 80% of IR gave the best water management in having both high K desorption from used soil and significant K uptake by corn plant. This study also concluded that a significant relationship was observed between the rate constants of the best fitted models and K uptake by corn gave trusting in such models in describing the uptake of this essential nutrient from sandy soil.***

Index Terms: **K adsorption, Kinetic models, Power Function equation**

I. INTRODUCTION

Numerous policies face a crisis of chemical fertilizers in Egypt between the policies of short-term and include market regulation, and policy management of the crisis, and policies to bridge the gap, also policies of the rationalization of use, and long-term policies include the establishment of new plants for fertilizers and the advancement of factories based and multi-ministerial decisions in this way shows how the depth of the crisis chemical fertilizers.

Of these types, potassium (K^+) is the third major essential nutrient for plant growth. It plays an essential role for enzyme activation, protein synthesis and photosynthesis, its importance in agriculture is well recognized [19]. Although the distribution of K^+ forms differs from soil to soil as a function of the dominant soil minerals present, total soil K^+ reserves are generally large. Soil K^+ is typically divided into four forms: soil solution K^+, exchangeable K^+, non-exchangeable K^+, and K^+ in soil minerals. Worth to mention that previous studies of the effect water scarcity on plant growth have shown that drought stress induces morphological, physiological and biochemical changes, including changes to photosynthesis, plant height, dry matter production, leaf area and grain yield [23]. However, the effects of drought on K desorption or bioavailability was little handled.

Reference [4] indicated that mass flow is the main driving force for the movement of Ca^{2+} and Mg^{2+}. In contrast to mass flow, diffusion is an important factor for K^+ and PO_4^{-3}. By using the rhizobox system, [20] found that the depletion of K^+ in the rhizosphere suggests that K^+ movement is governed by diffusion.

Kinetics of chemical reactions in soil and aquatic environments is a topic that is of extreme importance. There are dynamic equilibrium and kinetic reactions between the different forms of soil K that affect the level of soil solution K at any particular time, and thus, the amount of readily available K for plants. Levels of soil solution K are determined by the equilibria and kinetic reactions between the other forms of soil K [14]. Reference [21] developed a new rhizobox system to study the nutrient movement in the rhizosphere.

An array of kinetic methods has been used to measure the rates of soil chemical process. Anion exchange resin (AER) for studying the kinetics of soil chemical process was established by [3]. It is well known that in batch technique the suspension is agitated using reciprocating shaker for several time intervals, and then the suspension is usually centrifuged to separate a clear supernatant solution for subsequent analysis. The use of centrifugation to separate the liquid from solid phase has several disadvantages in having inaccurate results of ion concentrations [17]. Centrifugation would create electro kinetic effects close to soil constituent's surfaces that would alter the ion distribution [18].

Electrical stirred Flow Unit (ESFU) was manufactured in NRC by [22] to try to elevate the most common problems found in such techniques. The estimation of K^+ available to crops as well as most of the K^+ fertilizer recommendations are based on soil analyses and do not consider K^+ release from non- exchangeable fractions [12]. The common methods of estimation of K^+ available to crops as well as most of the K^+ fertilizer recommendations are based on soil analyses which mean a waste of time and a lot of efforts [12].

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Although there are a lot of articles related with K release, there is a scarcity interested with the effect of water regime on K bioavailability from the kinetic perspective.

The objectives of this work are to evaluate in short period the modified set up (ESFU) on the kinetics of potassium desorption from sandy soil fertilized with both compost, chemical fertilizers and mixture of both and the trust of used models in describing K desorption from the treated sandy soil. Also, this work focused on the effect of IR on K desorption will be also studied.

II. EXPERIMENTAL LOCATION

El-Nubaria site represents new reclaimed area (North of Delta) in Egypt and adjacent to a lot of farms cultivated with different crops. The cultivated area of National Research Centre Farm previously cultivated with different crops before our experimental started and used drip irrigation system as a main type of irrigation. This soil is characterized by soil acidity equal to 8.11, EC 1.32 dS/m., the soluble cations values were 0.48, 0.12, 0.69 and 0.06 meq/100g soil for Ca2+, Mg2+, K+ and Na+, respectively, 0.22, 0.77 and 0.36 meq/100g soil for HCO3-, 0.77 and SO4--, respectively. The texture of this soil was sandy loam. Available nutrient determined were 7.9 and 186.6 ppm for P and K, meanwhile OM and total CaCO3 values were 0.47 and 24.9%. All chemical characterizations of used were done as described by [11].

Experimental technique

Field experiments were conducted in El-Nubaria site to study the relationship between type of fertilizer applied and the kinetic rate of K desorption from the soil. Three techniques of fertilization were applied to Corn cultivar Single cross 129 (white) (*Zea mays* L.) was obtained from Ministry of Agriculture, Giza, Egypt, these techniques are represented by:

T1: 100% compost prepared by the authors, the rate of application based on the recommended rate of total nitrogen for corn crop

T2: 75% compost + 25% chemical fertilization

T3: 50% compost + 50% chemical fertilization

T4: 100% recommended chemical fertilization for corn applied based of the recommendation rate of the crop (120 kg N/fed. as ammonium sulfate + 30 kg P_2O_5/fed. as super phosphate + 24 kg K_2O/fed. as potassium sulfate)

C: Control, for this soil, the release of K based on the native K fertilizer applied in previous crops grown in the same used soil.

Organic material, phosphorous and potassium fertilizers were added before sowing. Nitrogen fertilizer was added in three equal portions before cultivation, after two weeks from cultivation and after three weeks from second addition, respectively. All amendments were manually spread. It is worth to mention that the experiment continued for two seasons.

Table (1): Some physical and chemical properties of the compost in two seasons.

Season	Total nutrients N%	P %	K %	OM %	C/N	EC dS/m 1:5	pH 1:2.5	OM %	WHC
First season	1.23	0.79	2.04	25.21	11.9	5.63	7.47	25.2	110
Second season	1.1	0.82	2.20	37.69	17.8	5.50	7.44	37.7	160

Agricultural practices were followed the recommendations of Ministry of Agriculture in Egypt. Enrichment compost with effective microorganisms (EM: *Bacillus subtilis* F.50, F.30, B. *Theremogensid* F.64, *Trichoderma reesei* F.418 and *Sacchromyces cerevisiae* F N.10) were used and prepared as described by [1]. EM was brought from the Biotechnology Unit, Microbial Chemistry Dept. N.R.C. Some physical and chemical properties of the compost in two seasons are shown in Table (1). Portions of dried maize plant materials were ground, wet-digested and analyzed for K as described by [7].

Kinetic Models

The linear forms of several kinetic models i.e. First- Order, parabolic Diffusion, Elovich and power function (modified Freundlich) as shown in table (2) were compared in their ability to fit the K desorption from compost-fertilized treated soil samples using nonlinear regression procedure [15]. The higher correlation coefficient of determination R2 and the lower standard error SE, the best-fitted equation(s) described the kinetics of K desorption in different treatments.

Irrigation Requirement IR

Since the irrigation requirement IR calculated for corn (Zea mays L.) is an important factor in K bioavailability in soil system, we evaluated two irrigation moisture regimes represent 80% and 60% of total water requirement on K desorbed from the used soil and subsequently the K bioavailability by corn. The data of IR was calculated by average 8 years of meteorological parameters using CROPWAT computer model [9] (according to the climatic data recorded at El-Bustan Weather Station in North of Delta, based on calculation of Penman Monteith equation and the Kc values presented in the program and also illustrated in [2].

Table (2) kinetic models applying to describe potassium

Model	Equation	References
Power Function	q=kd tb\	[5]
Elovich	Q= ln α + β ln t	[5]
First-order	Log q = log q0 – k1 t	[5]
	q = b + R t1/2	[22]

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adsorption in the studies soils

q = the amount of potassium adsorbed in time t
kd = desorption rate coefficient in mg kg^{-1}soil min^{-1}
b\ = intensity constant in mg K kg^{-1}soil
α = a constant related to the initial rate of K adsorbed in mg K kg-1min^{-1}
β = a constant in mg kg^{-1}soil
b = intensity constant in mg K kg^{-1}soil
R = the apparent diffusion rate coefficient in mg K kg^{-1}soil min^{-1}
q0 = the maximum amount of K adsorbed mg K kg^{-1}soil
k1= the rate constant of the reaction in sec-1
t = time (min).

Statistical analyses:

Experimental treatments were replicated three times in 2-Way Randomized Block design (2-WRB) with irrigation treatments in main plots and fertilizers treatments distributed randomize within main plots. All data obtained from this study were statistically analyzed through analysis of variance (ANOVA) and least significant difference (LSD) was applied to make comparisons among treatment means according to [10].

III. RESULTS AND DISCUSSION

Kinetic of K desorbed from the studied soil as affected by type of fertilizer applied Although data not shown, the authors didn't find any significance between the application of 100 and 80% of IR on K uptake by corn, consequently, the discussion will be only focused on 80 and 60% of IR. Figure 1 represents the kinetics of potassium release from the treated soils and irrigated with 80% of IR after 14 days of reaction time. Because of the S shape observed in entire reaction time for K desorption in different treatments, data was divided into two stages, the first stage from 1-60 min represents the rapid stage of K released and the second one 120-1880 min represents the rest of reaction time.

Dividing the entire reaction time into two stages, led to have almost straight lines in both short reaction time (A) and long one (B) for all treatment tested which represents that more than one mechanism controlled the release of K from the treated soils. In addition, it should be mention that all treatment kept their order in K release in both stages. It is generally believed that there is no single equation that described equally well the kinetic data of all soils (Ravan and Hosner, 1994).

Data showed that K release from compost individually gave the lowest values compared with other treatments tested except control, by mixing compost prepared with the chemical fertilizers, K released from the soil was significantly increased. Numerically, the maximum K desorbed from the soil was 221ppm in case of compost; gradually increase up to 257 and 277 ppm in T2 and T3, respectively. The highest value however, was observed in chemically fertilized soil 303 ppm.

Although data not shown, decreasing the irrigation requirement to 60% directly decreased the rate of K desorption in all treatments including control. For example, decreasing the IR led to decrease the maximum release of K in T4 to about 255 ppm which almost represents 15% less than the high IR applied. Also, in the organic treated soil T4, the decreasing order observed was less than the chemical treated one by about 200 ppm (about 10% less), worth to mention that the other treatments were decreased with varied percentage of decrease

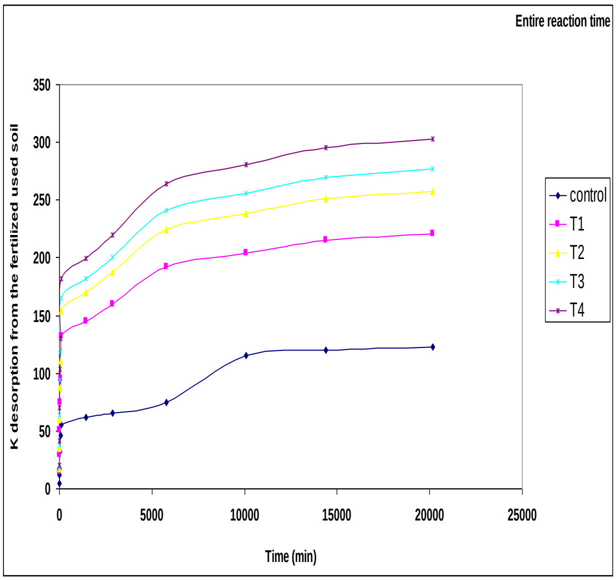
4.2 Kinetic parameters of K desorption from the fertilized used soil as affected by irrigation cycle and type of fertilizer applied Prior the narrative in kinetic work, it should be mention that selection of the best fitted model(s) was based on the higher coefficient of determination and the lower standard error the best fitted model(s). Accordingly, although the all models well described the kinetic data we found that MFE was the best in describing the rate of K desorption (R2 ranged between 0.98**-0.99**) since it gave the lower SE, followed by Elovich (0.96**-0.99**) and for less extent 1st order and parabolic diffusion models (R2 ranged between 0.89**-0.92**).

Through the entire reaction time, the succession of more than one model in describing the rate of K desorption in different treatments applied meaning that different mechanisms controlled K release from the treated sandy soil or in other words different forces retained K in treated soil and subsequently, the bioavailability of K in such systems.

The power function (modified Freundlich) equation in the linear form is: lnCt = ln kd + b\ ln t. The integrated form is qt = kd t b\

where qt the amount of K release at time t, kd and b\ are constants. Taking the derivation of integrated form:

dq / dt = kd b\ t b\ -1 .



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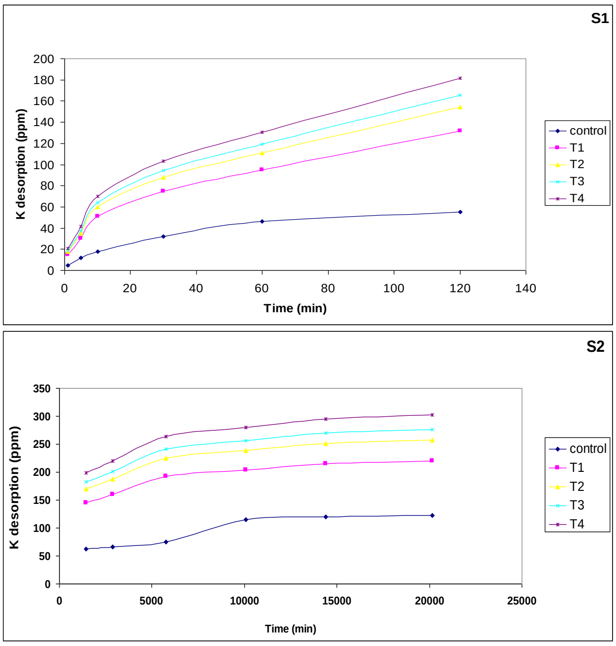

Fig (1) Kinetic of K desorbed from the treated soils in both entire reaction time and different stages.

where kd is directly proportional to the rate of K release and was considered as the apparent desorption rate coefficient. The effect of b\, the capacity factor, on potassium release is more complex since there are different soil parameters and treatments such as organic amendments controlled K reaction in soil system. The reaction rate is proportional to kd only at t = 1 in which case:

$$dq/dt = kd\, b\backslash .$$

The b\ value is convenient to use as an estimate of the initial release rate when comparisons are made between power function equations. It is, however, designated as the reversibly adsorbed potassium.

The kinetic parameters which describe the kinetic parameters of K release from soil irrigated with 80 and 60% of IR are presented in tables 3 and 4. The rate constants namely; Kd of modified Freundlich, R of the parabolic Diffusion, k1 of the first-order and β of Elovich equations, all were considerably increased in the 100% chemical fertilization applied T4 than in both purely organic fertilization and the mixture of chemical and organic types. The average values of MFE as an example in chemical fertilized soils were 0.26 decreased to 0.22 in organic fertilized soil and take in between values for the mixture of both organic and inorganic fertilizers. In contrast, data showed that the capacity factor b of MFE took a reverse trend reached to 1.55 for organic fertilizer and decreased to1.41 in chemically fertilized soil.

Decreasing the IR to 60% applied to corn decreased the Kd values for both chemical and organic fertilized soil to be 0.24 and 0.21 in chemical and organic fertilized soils respectively. Also, it was noted that although the capacity factor gave a reverse trend, the obtained values was less than the 80% IR.

The β value, the slop of the kinetic data plotted according to Elovich equation (ln of time against P concentration) was shown by [8] to be inversely proportional to the soil supplying power of ion to plant. Reference [5] showed that the decrease in 1/β and or increase in α enhance the reaction rate. Reference [16] reported that β constant is an important parameter to define desorption rate dI/dt throughout the whole dissolution period of added ion. Moreover, a low β value is associated with higher dissolution rate and a greater buffering of ion- dissolution rate with the increase in ion concentration, compared with highβ.

The β value in Elovich equation shown to be inversely related to the K supplying power of the soil was 8.25 and 8.8 in T2 and T3, respectively represented the preferring of applying 50:50 chemical to organic fertilizers instead of increasing organic fertilizer over the chemical one. These results clearly demonstrate the higher potential of K release in the soils treated with chemical ones. The constants that describe the capacity of K in used models i.e. b\ in modified Freundlich, b in parabolic diffusion, q0 in the first-order models, all were higher in the fertilized soil compared to control. Therefore plant K uptake from the treated soils was considerably higher than from the untreated one. In all cases, however, the kinetic parameters decreased by decreasing of IR applied.

Diffusion model and the mechanism were also succeeded in describing the rate of K desorbed from the treated soil. According to data presented in tables 3 and 4, decreasing the IR from 80% to 60 %, led to decrease the rate of K from 1.79 to 1.69 in T4 (chemically fertilized soil), in other words, the diffusion mechanism of K to plant used was influenced by the IR and assumed to be influenced the concentration of K in plant. The capacity factor represented by b decreased from 1.41 to 1.35. It is well known that the role of K in the plant could be summarized in photosynthesis and plant food formation, sugar and carbohydrate production, transport, and storage, however, the important function related with cell turgor and through this the opening and closing of leaf stoma. This in turn controls the plants ability to effectively respond to drought stress. Also, K is transported within the soil and is absorbed by plant roots in the soil water. Therefore a water deficiency results in less K absorption.

Correlation analysis between K desorption from the fertilized soils and the uptake of K by corn

The K uptake by corn grown in treated soils of the new reclaimed area was correlated with the constants of the four tested kinetic models (table 5). The correlation coefficient was highly significant with the kinetic parameters representing the rate of K release in the order: The modified Freundlich rate coefficient (Kd) (r = 0.98**) > the rate constant (β) in the Elovich equation (r = 0.97***) > the rate of K release coefficient (R) in the parabolic diffusion (r = 0.96**) > the initial rate of K-release (q0) in 1st order equation (r = 0.85**). The β coefficient in Elovich which was found to be inversely correlated with the K supplying power of the soil [13], and inversely related to the rate of K release

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from soil [6] was inversely correlated with the plant K uptake (r = 0.97**). It should be recalled that both Elovich and modified Freundlich equations offered the best fit to the K release data from the tested adequate fit to the description data, its rate of diffusion coefficient (R) gave the high correlation coefficient with the cumulative K uptake and with the uptake of K. Similarly the first-order equation offered a lower fit to the K release data but the rate constant k0 of the equation was highly correlated with plant K uptake.

The correlation coefficients in the same table show that K bioavailability may be less dependent on the capacity than the rate parameters due to lower correlation coefficients and the less significant correlation between the plant K uptake and the intensity parameters in the tested equations. The b\ value in modified Freundlich equation denotes the K capacity and was not significantly correlated with plant K uptake. The b value, which is the concentration of K at the time of ESFU work in the diffusion equation, showed significant correlation with plant K uptake but the correlation coefficient (r =0.63***) was considerably less than with the rate of K release (r = 0.95***). The q0 value in the first-order kinetic equation which represent the K intensity was less significantly correlated with K-uptake (r = 0.57**) than the rate kinetic R (r = 0.83***).

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Table (3) kinetic parameters of selected equation describe K desorption from compost-fertilized soil under application of 80% of IR.

Treat	Modified Frundlich equation		Parabolic diffusion equation		Elovich equation		1St. order equation	
	kd	b\	R	b	β	α	K1	q0
C	0.19	0.96	0.77	24.76	-4.76	11.43	-0.65	4.3
T1	0.22	1.55	1.30	92.34	7.00	28.97	-0.63	3.43
T2	0.23	1.51	1.52	84.25	8.25	26.46	-0.58	3.40
T3	0.24	1.48	1.63	78.58	8.80	24.65	-0.53	3.35
T4	0.26	1.41	1.79	67.13	9.69	21.08	-0.50	3.23

Treat.	MFE		Parabolic diffusion equation		Elovich equation		1St. order equation	
	kd	b\	R	b	β	α	K1	q0
C	0.17	0.93	0.73	21.56	-3.43	10.40	-0.86	1.43
T1	0.21	1.52	1.19	88.25	6.37	26.36	-0.82	3.12
T2	0.22	1.48	1.44	80.16	7.51	24.08	-0.71	2.92
T3	0.23	1.42	1.52	74.38	8.01	22.43	-0.63	2.85
T4	0.24	1.35	1.69	63.23	8.82	19.18	-0.50	2.79

Table (4) kinetic parameters of selected equation describe K desorption from compost-fertilized soil under application of 60% of IR.

The b value, which is the concentration of K at the time of ESFU work in the diffusion equation, showed significant correlation with plant K uptake but the correlation coefficient (r =0.63***) was considerably less than with the rate of K release (r = 0.95***). The q0 value in the first-order kinetic equation which represent the K intensity was less significantly correlated with K-uptake (r = 0.57**) than the rate kinetic R (r = 0.83***).

The high and significant correlation between K uptake and the kinetic parameters of the best models, namely kd, R, β, and k, however, gave higher correlation's with the plant K uptake indicating that these parameters, specially the diffusion rate coefficient (R) can better estimate K-bioavailability than the other methods in describing K uptake by plant. With an exception observed in the 1st order, Potassium uptake by corn in the 2nd season highly significant r values were observed with the capacity factor more than the rate parameters of tested models. For example, in the MFE, the r value of b\ was 0.91** meanwhile the Kd was 0.72ns, this result may represents the variation of K mechanism with repeating of crop cultivated in the same soil. Worth to mention that same trend was observed in diffusion and Elovich model.

The α parameters in Elovich equation that may stand for the initial rate of K release had no significant relation with plant K uptake. As with the 1st season, both the b value in the diffusion equation and the q0 in the first-order equation, which represent K capacity, showed high significant correlation with plant K uptake(r = 0.94** and 0.97**).

Almost all kinetic parameters describing the rate of K release in different treatments of the studied soil were highly correlated with plant K uptake and may be used as indices for K bioavailability if the rate of release proved to be a determining step for K-uptake by plant. In the 2nd season, however, only the 1st order rate coefficient (k1) was highly correlated with plant K-uptake. This parameter assumed superiority over other capacity parameters possible indices for K bioavailability in the 1st season. These results agree with those of [5], [6], [13] and [16] and. Their results on phosphate showed that both constants were related to the P supplying power of soil and or the rate of P release from soil had significant relationship between β and P bioavailability to sorghum did not establish that P dissolution is a rate determining step to P-uptake since similar relationship existed between β and equilibrium ion. On the other hand, [6] argued that the potential rate of P release from the soil as a whole was at least 250 times as great as the rate of P-uptake by crop.

Our results indicate that the rate of K release from the soil consistently influenced by irrigation regime applied to soil system. In addition, these intensity parameters could be used as indices for K bioavailability since they showed superiority compared to capacity parameters. The kinetic method used to extract K i.e. ESFU was offered a simple and accurate method for K release. In addition, the important of this study may emphasized that the 80% of IR was the best irrigation treatment in having optimum condition of K desorption to growing plants since it makes stimulation of plants to increased K uptake. Decreasing IR to 60%, led to decrease the rate parameters and the ability of the tested crop to absorb K.

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Table (5) Correlation coefficient (r) for the relationship between K uptake by Corn plants grown in treated soil and kinetic parameters of selected equations.

seasons	MFE		Diffusion		Elovich		1st order	
	kd	b\	R	B	β	α	K1	q0
Uptake 1st season	0.98**	0.89**	0.96**	0.94**	0.97**	0.96**	-0.85**	0.97**
Uptake 2nd season	0.72ns	0.91**	0.81*	0.86**	0.77ns	0.83*	-0.92**	0.78ns

ns= not significant (p <0.01)